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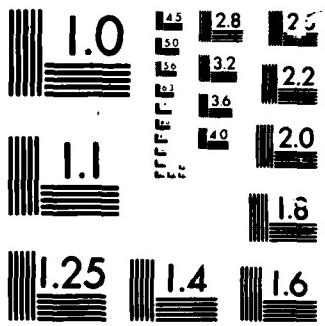
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Synthesis of Some Arsinogallanes and the Novel
Rearrangement of a Dimeric Bisarsinogallane,
Bis{bis[bis(trimethylsilylmethyl)arsino]chlorogallane}

by

C. G. Pitt, A. P. Purdy, K. T. Higa, and R. L. Wells

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Department of Chemistry
Durham, NC 27706

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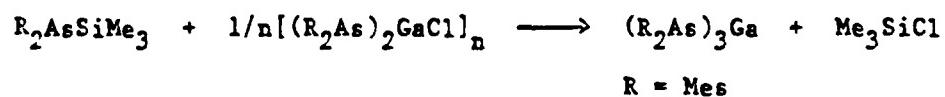
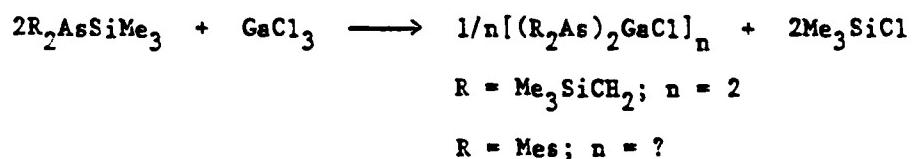
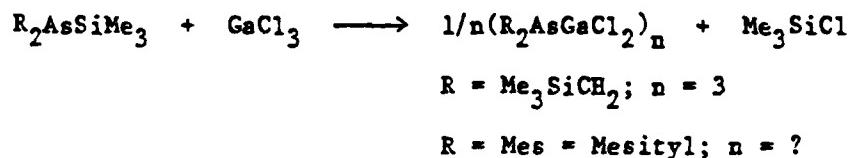
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Synthesis of Some Arsinogallanes and the Novel
Rearrangement of a Dimeric Bisarsinogallane,
Bis{bis[bis(trimethylsilylmethyl)arsino]chlorogallane}

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Abstract: Silylarsines have been used to synthesize arsinogallanes according
to the following equations.



Dynamic NMR spectroscopy studies of one of these, $\{(Me_3SiCH_2)_2As\}_2GaCl\}_2$,
show it has novel fluxional properties. Changes of the ¹H and ¹³C NMR spectra
were reconciled by rapid exchange of the endo- and exocyclic

Thermal reactivity (Delta G) +01-

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bis(trimethylsilylmethyl)arsino groups, and spectra simulated on this basis were in excellent agreement with the experimental data. A value of 17.0 ± 0.0 kcal/mol for the ^{13}C exchange was obtained by line shape analysis, and was in agreement with the estimate derived from the coalescence temperature. The ^1H exchange process between 22 and 57°C could be simulated using the same kinetic parameters, without invoking pyramidal inversion of the exocyclic arsenics.

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Summary: Silylarsines, $R_2AsSiMe_3$ ($R = Me_3SiCH_2$ and Mesityl), have been used to synthesize arsimogallanes. Dynamic NMR spectroscopy studies of one of these, $\{(Me_3SiCH_2)_2As\}_2GaCl_2$, show it has novel fluxional properties.

Despite the importance of gallium arsenide in emerging semiconductor technology¹, the chemistry of this elemental combination had, until recently², been restricted to the studies of Coates *et al.*³ in the 1960's. They prepared two monoarsinogallanes by cleavage of one Ga-C bond in Me_3Ga with secondary arsines, resulting in the elimination of methane (Eq. 1). In our hands, this

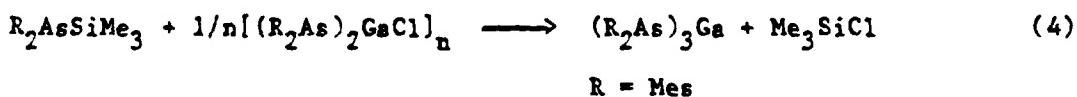
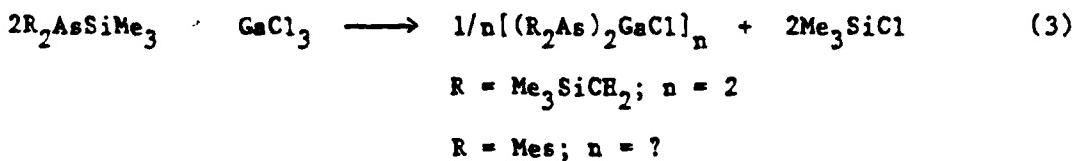
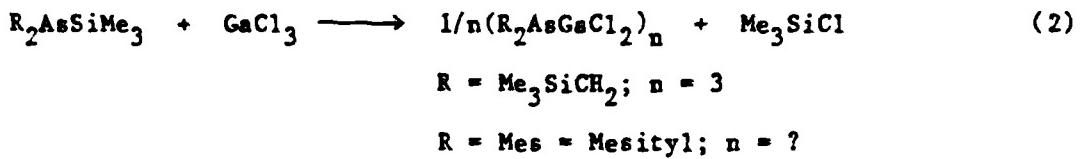


$R = Me; n = 3$

$R = Ph; n = 2$

method became impractical, and eventually failed, as the steric bulk of the substituents was increased;⁴ it was also found to be limited to the cleavage of a single Ga-C bond, and bis- and trisarsinogallanes, $[(R_2As)_{3-x}GaR_x]_n$

($x = 0, 1$), were inaccessible by this route⁵. A number of reports of the synthesis of main group compounds by metathetical elimination of trimethylchlorosilane⁶ prompted us to evaluate this route to arsinogallanes. Here we report its successful application (Eqs. 2-4), as well as the novel fluxional properties of a dimeric bisarsinogallane.



Silylarsines⁷ are readily available from lithium arsenides and trimethylchlorosilane, and their reaction with trichlorogallane proceeds stepwise in hydrocarbon solvents⁸. The only by-product, trimethylchlorosilane, is easily removed in vacuo, a considerable advantage when handling these air and moisture sensitive materials. This method did not yield $[(Me_3SiCH_2)_2As]_3Ga$; however, it did afford the sterically hindered trisarsinogallane $(Mes_2As)_3Ga$ (Eq. 4) which was first prepared by us from lithium dimesitylarsenide and trichlorogallane and shown to be monomeric by single-crystal x-ray analysis^{2d}. Lack of suitable crystals and the very low

solubility of the mono- and bisarsinogallanes containing mesityl groups (Eqs. 2 and 3), also prepared from lithium dimesitylarsenide and trichlorogallane)^{2d}, prohibited determination of their degree of oligomerization.

The bisarsinogallane $[(\text{Me}_3\text{SiCH}_2)_2\text{As}]_2\text{GaCl}$ was shown by cryoscopic measurements to exist predominantly as the dimer (Eq. 3) in benzene. Its ^1H (300 MHz) NMR spectrum in toluene at 22 °C, reproduced in Figure 1a, consisted of singlets at δ 0.25 and 0.27 (72H, Me_3Si), and 1.65 (8H, endo- CH_2As), and doublets (AB quartet, $^2J_{\text{HH}} = 13.5$ Hz) at 0.94 and 1.76 (8H, exo- CH_2As). The ^{13}C NMR spectrum at 21 °C consisted of four singlets at δ 0.05 and 1.15 (Me_3Si), 7.87 and 8.78 (SiCH_2As). Only a single isomer was detected, assigned the trans structure on steric grounds. The ^1H NMR spectrum was measured using a range of concentrations (5.9-111 mmolar) in an attempt to detect equilibria with monomeric or other oligomeric species; no significant changes in peak intensities of the dimer were observed. However, the relative intensities of the peaks due to trace impurities, $(\text{Me}_3\text{SiCH}_2)_2\text{AsH}$ and $(\text{Me}_3\text{SiCH}_2)_2\text{AsAs}(\text{CH}_2\text{SiMe}_3)_2$, resulting from slight hydrolysis and decomposition of the dimer, respectively⁹, increased somewhat with decreasing concentrations.

On raising the temperature from 22 °C to 97 °C, the ^1H AB quartet collapsed to a broadened doublet which then approached coalescence with the singlet. Likewise, the ^{13}C methyl peaks broadened and coalesced at 78 °C (Fig. 1b). This spectroscopic behavior was reversible, with the exception of the formation of a small amount of the above diarsine. These changes of ^1H and ^{13}C spectra, when coupled with the dimeric molecular weight and the absence of other species in the spectra, can only be reconciled by rapid exchange of the

endo- and exocyclic bis(trimethylsilylmethyl)arsino groups. Spectra simulated on this basis were in excellent agreement with the experimental data. A ΔG^\ddagger of 17.0 ± 0.0 kcal/mol for the ^{13}C exchange was obtained by line shape analysis¹⁰, and was in agreement with the estimate derived from the coalescence temperature¹¹. The ^1H exchange process between 22 and 57 °C could be simulated using the same kinetic parameters, without invoking pyramidal inversion of the exocyclic arsenics. Evidently, the latter is not a significant process in this temperature range, despite the fact that the electropositive gallium might be expected to reduce the inversion barrier.

The equivalency of the bis(trimethylsilylmethyl)arsino groups can be achieved by associative or dissociative mechanisms. The associative mechanism is intuitively more attractive, involving an intermediate in which one or both gallium atoms assume a five-coordinate square pyramidal configuration (Fig. 2). There is ample precedent for five-coordinate gallium.¹² The dissociative mechanism requires cleavage of one endocyclic Ga-As bond, then reclosure of the ring, in a trans-cis-trans equilibrium process. Dissociation into two monomers is considered energetically improbable.

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References and Notes

- (1) Keyes, R.W. Science 1985, 230, 138.
- (2) (a) Wells, R.L.; Purdy, A.P.; McPhail A.T.; Pitt, C.G. Abstracts of Papers, 189th National Meeting of the American Chemical Society, Miami Beach, FL, April 1985; American Chemical Society, Washington, D.C., 1985, INOR 26.
(b) Wells, R.L.; Purdy, A.P.; McPhail, A.T.; Pitt, C.G. J. Chem. Soc., Chem. Commun. 1986, in press. (c) Wells, R.L.; Purdy A.P.; McPhail, A.T.; Pitt, C.G. J. Organomet. Chem., submitted. (d) Pitt, C.G.; Higa K.T.; McPhail A.T.; Wells, R.L. Inorg. Chem., submitted.
- (3) (a) Coates, G.E.; Graham, J.J. J. Chem. Soc. 1963, 233. (b) Beachley, O.T.; Coates, G.E. J. Chem. Soc. 1965, 3241.
- (4) For example, $[(\text{Me}_3\text{SiCH}_2)_2\text{AsGaPh}_2]_2$ was prepared by this method (ref. 2c); however, heating Ph_2AsH and $(\text{Me}_3\text{SiCH}_2)_3\text{Ga}$ (1:1 molar ratio) at 60°C for 2 days, afforded impure $[(\text{Ph}_2\text{AsGa}(\text{CH}_2\text{SiMe}_3)_2)]_n$, $\text{Ph}_2\text{AsAsPh}_2$, Me_4Si , H_2 , and unreacted $(\text{Me}_3\text{SiCH}_2)_3\text{Ga}$; no evidence for the desired product was obtained when $(\text{Me}_3\text{SiCH}_2)_2\text{AsH}$ and $(\text{Me}_3\text{SiCH}_2)_3\text{Ga}$ (1:1 molar ratio) were heated at $100\text{-}120^\circ\text{C}$ for 1 month.
- (5) For example, when Ph_2AsH and Me_3Ga (2:1 molar ratio) were heated at $80\text{-}100^\circ\text{C}$ for 2 days, no evidence for any product other than CH_4 , $(\text{Ph}_2\text{AsGaMe}_2)_n$, and unchanged Ph_2AsH was obtained.
- (6) See for example: (a) Abel, E.W.; Armitage, D.A.; Willey, G.R. J. Chem. Soc. 1965, 57. (b) Wells, R.L.; Collins, A.L. Inorg. Chem. 1966, 5, 1327. (c) Abel, E.W.; Illingsworth, S.M. J. Chem. Soc. (A) 1969, 1094. (d) Harman, J.S.; McCartney, M.E.; Sharp, D.W.A. J. Chem. Soc. (A) 1971, 1547. (e) Goetze, R.; Noeth. H. Z. Naturforsch. 1975, 30b, 875. (f) Nutt, W.R.;

Stimson, R. E.; Leopold, M.F.; Rubin, B.H. Inorg. Chem. 1982, 21, 1909. (g)
Hoffmann, G.G. J. Organomet. Chem. 1984, 273, 187.

(7) See for example: (a) Abel, E.W.; Honigschmidt-Grossich, R.; Illingworth, S. M. J. Chem. Soc. (A) 1968, 2623. (b) Abel, E. W; Crow, J.P. J. Organomet. Chem. 1969, 17, 337. (c) Mislow, K.; Senkler, G.H.; Cook, R.J.; Casey, J.P.; Baechler, R. D. J. Am. Chem. Soc. 1972, 94, 2859. The new silylarsines used in this work were prepared under strict inert atmospheric conditions as follows. $(\text{Me}_3\text{SiCH}_2)_2\text{AsSiMe}_3$. A mixture of $(\text{Me}_3\text{SiCH}_2)_2\text{AsH}^{2c}$ (6.0 g, 24 mmol) and n-BuLi (10 mL, 2.4 M in hexane) was stirred at 60 °C for 1 day under nitrogen pressure and then cooled to 0 °C, before adding dropwise excess Me_3SiCl (4 mL in 20 mL pentane). After refluxing for 0.5 h, addition of 10 mL of Et_2O , and standing at room temperature for 1 day, the mixture was filtered; following solvent removal, distillation afforded the silylarsine as a colorless liquid (5.8 g, 75% yield), bp 40 °C/0.04-0.35 Torr. Anal. Calcd (Found) for $\text{C}_{11}\text{H}_{31}\text{AsSi}_3$: C, 40.96 (40.83); H, 9.69 (9.82). NMR: ^1H (80 MHz) (C_6D_6) δ 0.80 and 0.46 [AB pattern ($^2J_{\text{HH}} = 13.6$ Hz) 4H, CH_2], 0.18 (s, 9H, AsSiMe_3), 0.15 (s, 18H, CSiMe_3); ^{13}C (22.5 MHz) (C_6D_6) δ 5.7 (s, CH_2), -1.5 (s, AsSiMe_3), 0.37 (s, CSiMe_3). $\text{Mes}_2\text{AsSiMe}_3$. (a) Mesitylmagnesium bromide (274 mmol) in THF (200 mL) was added over a 45 min period, with stirring, to AsCl_3 (25.0 g, 138 mmol) in THF; after 12 h, degassed water (125 mL) and conc HCl (60 mL) were added, and the mixture was stirred for 1 day. Following extraction with benzene, the volatiles were removed by distillation leaving crude, solid Mes_2AsCl (45.1 g, 94% yield). NMR: ^1H (80 MHz) (C_6D_6) δ 6.60 (s, 4H, ring CH), 2.35 (s, 12H, Me), 2.02 (s, 6H Me). (b) The crude Mes_2AsCl was dissolved in THF (500 mL) and LiAlH_4 (10.0 g, 263 mmole) was added. After stirring the solution for 46 h, addition of water (220 mL), extraction with

benzene, and removal of volatiles, distillation afforded Mes_2AsH as a colorless liquid (14.6 g, 58% yield), bp 145–150 °C/0.01 Torr. Anal. Calcd (Found) for $\text{C}_{18}\text{H}_{23}\text{As}$: C, 68.79 (68.66); H, 7.38 (7.55). NMR: ^1H (80 MHz) (C_6D_6) δ 6.70 (s, 4H ring CH), 4.83 (s, 1H, AsH), 2.27 (s, 12H, Me), 2.07 (s, 6H, Me). IR: (Nujol) 2130 cm^{-1} (s, AsH). (c) A mixture of Mes_2AsH (14.62 g, 46.42 mmol), $n\text{-BuLi}$ (20 mL, 2.4 M in hexane), and ligroin (150 mL) was heated at 50 °C for 20 min. Filtration gave solid, yellow Mes_2AsLi which was dried under vacuum (14.41 g, 96.7% yield). Anal. Calcd (Found) for $\text{C}_{18}\text{H}_{22}\text{AsLi}$: C, 67.51 (67.15); H, 7.82 (7.95). NMR: ^1H (80 MHz) (C_6D_6 + 1 drop THF) δ 6.93 (s, 4H, ring CH), 2.61 (s, 12H, Me) 2.29 (s, 6H, Me). (d) A mixture of Mes_2AsLi (3.79 g, 11.8 mmol), Me_3SiCl (10 mL), and ligroin (100 mL) was heated to reflux, stirred at room temperature for 12 h, and filtered. The filtrate was concentrated in vacuo to an oil which eventually crystallized as colorless $\text{Mes}_2\text{AsSiMe}_3$ (4.46 g, 98% yield), mp 62–66 °C. Anal. Calcd (Found) for $\text{C}_{21}\text{H}_{31}\text{AsSi}$: C, 65.26 (65.53); H, 8.08 (8.24). NMR: ^1H (80 MHz) (C_6D_6) δ 6.74 (s, 4H, ring CH), 2.31 (s, 12H, Me), 2.10 (s, 6H, Me), 0.28 (s, 9H, SiMe_3).

(8) The arsinogallanes were prepared under strict inert atmospheric conditions as follows. $[(\text{Me}_3\text{SiCH}_2)_2\text{AsGaCl}_2]_3$. A solution of $(\text{Me}_3\text{SiCH}_2)_2\text{AsSiMe}_3$ (0.5 g, 1.55 mmol) in pentane (6 mL) was added to a solution of GaCl_3 (0.27 g, 1.55 mmol) in pentane (6 mL). After several min, a precipitate formed. Removal of the liquid phase, washing with pentane, and drying in vacuo gave the arsinogallane as a white solid (0.40 g, 67% yield), mp 170–184 °C (decomp.). Anal. Calcd (Found) for $\text{C}_{24}\text{H}_{66}\text{As}_3\text{Cl}_6\text{Ga}_3\text{Si}_6$: C, 24.64 (24.82); H, 5.68 (5.97); Cl, 18.18 (17.84); mol wt (cryoscopic in benzene), 1170 (1070 80); NMR spectrum shows one species. NMR: ^1H (250 MHz) (C_6D_6) δ 1.78 (s, 12H, CH_2), 0.23 (s, 54H, SiMe_3); ^{13}C (22.5 MHz) (C_6D_6) δ 7.4 (s, CH_2)

1.2 (s, SiMe₃). {[(Me₃SiCH₂)₂As]₂GaCl}₂. A solution of (Me₃SiCH₂)₂AsSiMe₃ (1.0 g, 3.10 mmol) in pentane (10 mL) was added to a solution of GaCl₃ (0.27 g, 1.55 mmol) in benzene (10 mL) and, after 1 day, the mixture was brought to a boil. Removal of the volatiles under vacuum gave a white solid that on slow recrystallization from ligroin (bp 95-110) afforded the bisarsinogallane as colorless crystals that were dried in vacuo (0.67 g, 80% yield), mp 101-128 °C (decomp.). Anal. Calcd (Found) for C₃₂H₈₈As₄Cl₂Ga₂Si₈: C, 31.82 (31.66); H, 7.34 (7.23); mol wt (cryoscopic in cyclohexane), 1207 (1100 ± 55). (Mes₂AsGaCl₂)_n^{2d}. A solution of Mes₂AsSiMe₃ (0.39 g, 1.0 mmol) in pentane (3 mL) was added to a solution of GaCl₃ (0.24 g, 1.36 mmol) in pentane (5 mL) and, after 5 min, the volatiles were removed. Washing the remaining solid with ligroin and drying in vacuo gave the arsinogallane as a white solid (0.46 g, 100% yield), mp 145-146 °C (decomp.). Anal.^{2d} Calcd (Found) for C₁₈H₂₂AsCl₂Ga: C, 47.63 (47.91); H, 4.89 (5.07). NMR: ¹H (80 MHz) (C₆D₆ + 2 drops THF) δ 6.69 (s, 4H, ring CH), 2.45 (s, 12H, Me), 2.06 (s, 6H, Me). [(Mes₂As)₂GaCl]_n^{2d}. Addition of a solution of Mes₂AsSiMe₃ (0.82 g, 2.1 mmol) in benzene (10 mL) to a solution of GaCl₃ (0.35 g, 2.0 mmol) in benzene (10 mL) gave a white precipitate of (Mes₂AsGaCl₂)_n. Addition of more Mes₂AsSiMe₃ (0.75 g, 1.9 mmol) resulted in a yellow-orange solution. Removal of volatiles after 3 h gave a yellow solid that on recrystallization from ligroin/benzene (1:1) afforded the bisarsinogallane as a yellow solid (0.35 g, 24% yield), mp 120-125 °C (decomp.). Anal. Calcd (Found) for C₃₆H₄₄As₂ClGa: C, 59.09 (59.27); H, 6.06 (6.02). NMR: ¹H (80 MHz) (C₆D₆) δ 6.68 (s, 8H, ring CH), 2.31 (s, 24H, Me), 2.06 (s, 12H, Me). (Mes₂As)₃Ga.^{2d} Addition of a solution of Mes₂AsSiMe₃ (0.67 g, 1.7 mmol) in benzene (10 mL) to a solution of GaCl₃ (0.15 g, 0.85 mmol) in benzene (10 mL) gave a yellow solution. Removal

of volatiles after 3 days gave a yellow solid which was shown by NMR to be an impure mixture of mono- and bisarsinogallane. Addition of more $\text{Mes}_2\text{AsSiMe}_3$ (0.21 g, 0.54 mmol) dissolved in benzene (10 mL) gave a yellow solid and orange solution. Filtration and removal of volatiles gave an orange oil to which ligroin (10 mL) was added. The resulting mixture was heated and allowed to cool; after 1 day the trisarsinogallane precipitated as an orange solid (0.11 g, 13% yield), mp 130–160 °C (decomp). Anal. ^{2d} Calcd (Found) for $\text{C}_{54}\text{H}_{66}\text{As}_3\text{Ga}$: C, 64.24 (64.55); H, 6.59 (6.59). NMR: ¹H (80 MHz) (C_6H_6) δ 6.63 (s, 12H, ring CH), 2.32 (s, 36H, Me) 2.05 (s, 18H, Me).

(9) Based on NMR, same peaks as in spectra of authentic samples of the arsine^{2c} and the diarsine. The diarsine was prepared by the reduction of $(\text{Me}_3\text{SiCH}_2)_2\text{AsCl}^{2c}$, ¹³ with zinc amalgam in refluxing methanol. Anal. Calcd (Found) for $\text{C}_{16}\text{H}_{44}\text{As}_2\text{Si}_4$: C, 38.54 (38.11); H, 8.89 (8.61). NMR: ¹H (80 MHz) (C_6D_6) δ 0.75 and 0.61 [AB pattern (²J_{HH} = 13.6), 8H, CH_2], 0.12 (s, 36H, SiMe_3); ¹³C (22.5 MHz) (C_6D_6) δ 10.2 (s, CH_2), 0.5 (s, SiMe_3).

(10) (a) The ¹³C (75.429 MHz) Me_3Si resonances from spectra at eleven temperatures (304, 309, 314, 319, 325, 330, 349, 350, 351, 352, and 365 °K) were compared with the simulations from the modified DNMR3^{10b} program, using Δν = 83.5 Hz, and a natural line width of 1.0 Hz, and the rate constants (5, 5, 11, 14.5, 29, 35, 150, 170, 185, 195, and 500 sec⁻¹) were obtained. A plot and least-squares analysis of ln (k/T) versus 1/T yielded a slope (m = -8295 ± 218) and an intercept (b = 22.98 ± 0.66). The equations^{10c} $\Delta H^\ddagger = -mR$, $\Delta S^\ddagger = R(b - 23.76)$, and $\Delta G^\ddagger = \Delta H^\ddagger - T\Delta S^\ddagger$ were used to obtain ΔH^\ddagger (16.5 ± 0.4 kcal/mole), ΔS^\ddagger (-1.55 ± 1.3 cal/mol·°K), and ΔG^\ddagger . (b) Binsch, G.; Kleier, D.A. The Computation of Complex Exchange-Broadened NMR Spectra, Program 165, Quantum Chemistry Program Exchange, Indiana University, 1970; revised by

Bushweller, C.H.; Bhat, G.; Letendre, L.J.; Brunelle, J.A.; Bilofsky, H.S.; Ruben, H.; Templeton, D.H.; Zalkin, A. J. Am. Chem. Soc. 1975, 97, 65; and by Caves, T. (North Carolina State University). (c) Sandstrom, J. "Dynamic NMR Spectroscopy"; Academic Press Inc.: London, England, 1982.

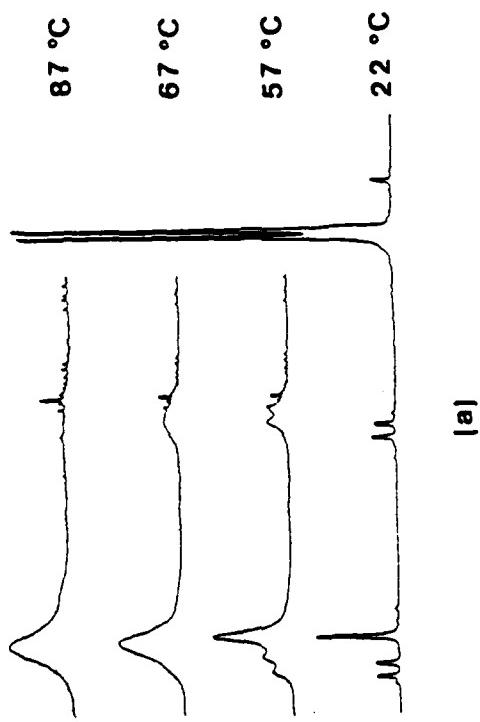
(11) $\Delta G_c^\ddagger = 17.0 \pm 0.1 \text{ kcal/mol}$; obtained (in cal/mol) from the equation $\Delta G_c^\ddagger = T_c [45.67 + 4.58 \log (T_c / \Delta v)]$, where $T_c = 351 \pm 1 \text{ }^{\circ}\text{K}$, and $\Delta v = 83.3 \text{ } 0.2 \text{ Hz}$.

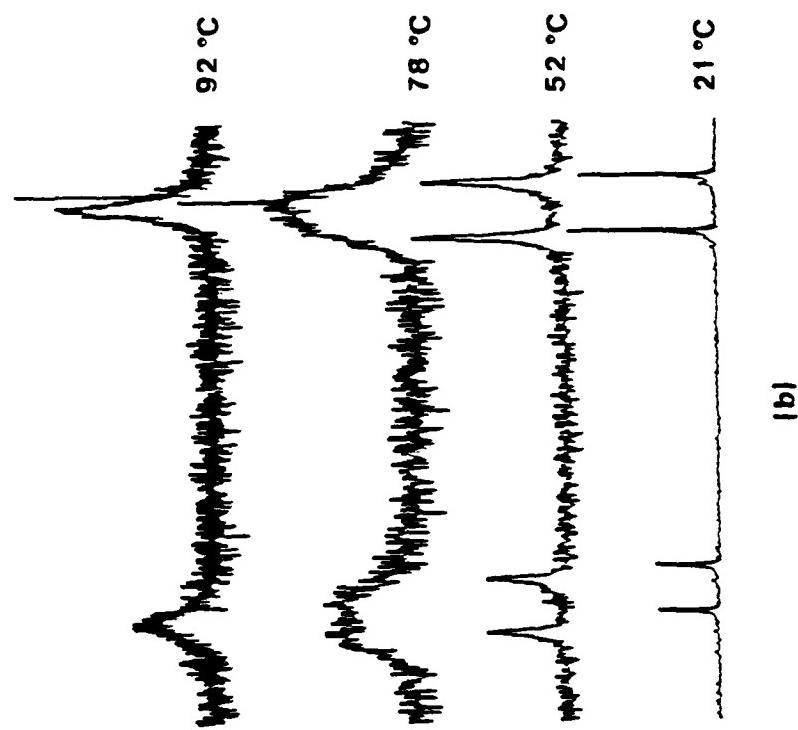
(12) See for example: (a) Pattison, I.; Wade K. J. Chem. Soc. (A) 1968, 2618. (b) Dymock, K.; Palenik, G.J. J. Chem. Soc., Chem. Commun. 1973, 884. (c) Rettig, S.J.; Storr, A.; Trotter, J. Can. J. Chem. 1975, 53, 753. (d) Rettig, S.J.; Storr, A.; Trotter, J. Can. J. Chem. 1976, 54, 1278. (e) McPhail, A.T.; Miller R.W.; Pitt, C.G.; Gupta, G.; Srivastava, S.C. J. Chem. Soc., Dalton Trans. 1976, 1657.

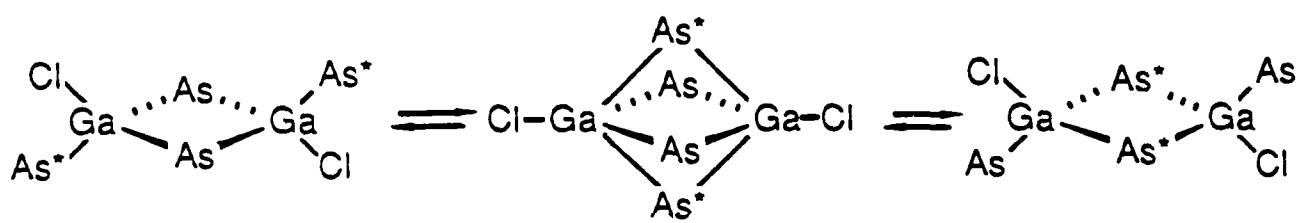
(13) Seyferth, D. U. S. Patent 2964550 1960.

Figure 1. (a) ^1H NMR Spectra and (b) ^{13}C NMR Spectra
of $\{[(\text{Me}_3\text{SiCH}_2)_2\text{As}]_2\text{GaCl}\}_2$.

Figure 2. Exchange of the Endo- and Exocyclic Arsino Groups.







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